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酸性物质对苏打盐碱土改良的研究进展*

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摘要: 苏打盐碱土作为我国后备土地资源, 具有极大利用潜力, 若加以合理的开发利用, 将会产生巨大收益。在苏打盐碱土改良中, 化学方法是一种重要的手段, 主要包括含钙制剂、有机类改良剂和酸性物质, 其中含钙制剂和有机类改良剂主要通过添加外源钙置换土壤胶体吸附的交换性钠和改善土壤结构促进水分淋洗来改良苏打盐碱土。苏打盐碱土自身富含碳酸钙, 以酸性物质作为化学改良剂可发生水解产生氢离子, 降低盐碱土 pH, 有利于溶解土壤中的碳酸钙, 为代换土壤中的交换性钠提供钙源, 从而减少外源钙质材料添加, 降低改良成本。本文通过综合分析国内外酸性物质对苏打盐碱土物理性质、盐碱性、养分利用和作物产量等方面的研究, 对酸性物质应用于苏打盐碱土改良的机理与实践进行总结, 并展望其未来发展趋势, 以期对苏打盐碱土治理与农业利用提供参考。

关键词: 苏打盐碱土; 酸性物质; 改良机理; 理化性质

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Progress of research on the improvement of saline-sodic soil using acidic substances*

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Abstract: As a reserve land resource in China, saline-sodic soil has great utilization potential; if properly developed and utilized, it will produce great benefits. Chemical methods are important for improving saline-sodic soils. In the early days, calcium-containing preparations, such as gypsum and phosphogypsum, were used to replace the exchangeable sodium adsorbed on soil colloids by adding exogenous calcium. However, saline-sodic soils are rich in calcium carbonate. Acid substances, as one of the chemical amendments, can be hydrolyzed to produce hydrogen ions. This reduces the pH of saline soil, which is conducive to dissolving calcium carbonate in soil and providing calcium sources for replacing exchangeable sodium in soil; thus reducing the addition of exogenous calcium materials and reducing the improvement cost. Based on a comprehensive analysis of domestic and foreign studies on the physical properties, saline properties, nutrient utilization, and crop yield of saline-sodic soil, this paper summarized the improvement mechanism and practice of the application of acid substances to saline-sodic soil, and forecasted its future development trend to provide a reference

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for saline-sodic soil management and agricultural utilization.

Keywords: Saline-sodic soil; Acid substances; Improvement mechanism; Physical and chemical properties

据联合国教科文组织的不完全统计,全球范围内约有 $8.3 \times 10^8 \text{ hm}^2$ 的盐碱土,约占全球陆地面积的 10%,广泛分布于世界各地^[1-2]。盐碱土是指包括盐土和碱土在内的各种盐化、碱化的土壤总称。盐土指含盐量多在 0.6%~1.0% 或者更高,碱土指碱化度(ESP)超过 15%,pH 高达 8 以上的土壤类型,其土壤表层盐分一般在 0.5% 以下^[3]。土壤盐碱可能通过土壤颗粒的分散对植物生长产生负面影响,破坏土壤结构和堵塞孔隙^[4-6],导致营养缺乏或失衡,即过量的 Na^+ 与 Ca^{2+} 、 K^+ 和其他阳离子竞争,降低它们的有效性^[7-8],产生渗透胁迫和离子毒害,影响植物生长。土壤盐碱已成为制约全球农业、畜牧业和经济发展的障碍。近年来,许多国家都十分关注这一问题,并对盐碱土的形成机理、水盐运移特点及盐碱土改良进行大量研究,以期找到改良盐碱土最佳方案。

中国土壤盐碱化程度较高,现共有盐碱地 $9.9 \times 10^7 \text{ hm}^2$,其中潜在盐碱化土壤约有 $1.7 \times 10^7 \text{ hm}^2$ ^[9],主要分布在东北、西北、华北和一些沿海平原地区。随着我国人口持续增长和经济快速发展,土地利用与耕地保护之间的矛盾日益突出,需尽快提高现有土地生产力和开发新的土地资源。合理改良利用盐碱地,对提高我国耕地面积,维护生态平衡,促进经济、社会、生态可持续发展具有重要意义^[10]。东北松嫩平原是我国盐碱土的主要分布区之一,该区土壤盐分主要以碳酸钠和碳酸氢钠为主,苏打盐碱土面积约占该地区盐碱土总面积的 70%^[11-13]。松嫩平原生态环境脆弱,自然因素和人为因素对生态环境造成的影响较大,其中土壤盐碱化问题最为突出^[14]。因此,对苏打盐碱土进行改良,无论从理论上还是实践中都具有重要意义。目前,苏打盐碱土的改良方法很多,效果显著,但其中一些方法由于条件限制和成本过高,难以大规模推广。寻求快速、高效、低成本、操作简便、可持续的苏打盐碱土改良技术成为当前的首要任务^[15]。

化学改良是苏打盐碱土改良的重要手段。化学改良主要通过离子交换和化学作用降低土壤交换性 Na^+ 的饱和度和盐碱度,再通过排水洗盐方法使盐分随水转移,促进土壤脱盐,改善土壤理化性质以及中和土壤碱度,减轻土壤盐碱化对作物的危害,增加作物生长所需的养分,达到改良盐碱土的目的^[14,16]。目

前改良苏打盐碱土化学改良剂主要有 3 类: 1) 含钙物质,如石膏、磷石膏等,主要以 Ca^{2+} 代换土壤胶体上的 Na^+ 为改良机理; 2) 酸性物质,如硫酸及其酸性盐类、磷酸及其酸性盐类,主要以中和碱及活化钙为改良机理,为替换可交换性 Na^+ 提供必要物质条件^[17]; 3) 有机类改良剂,如传统的腐殖质类(草炭、风化煤、绿肥、有机物料)、工业合成改良剂(禾康、聚丙烯酸等)、工业脚料糖醛渣等,主要以改善土壤结构促进水分淋洗为改良机理。

近年来,国内外学者对酸性物质改善土壤理化性质、提高土壤肥力、提高作物产量等问题进行深入研究。相关研究表明^[18-19]: 酸性物质能够促进土壤大团聚体形成,孔隙度增加,改善土壤结构,提升土壤稳定性。酸性物质也可通过水解作用释放 H^+ 降低土壤 pH,以便溶解土壤中的沉积钙,增加 Ca^{2+} 的含量,使 Na^+ 从土壤胶体置换到土壤溶液中,便于 Na^+ 随水分排出,达到改良土壤的效果^[20-21]。同时,可以提高土壤含水量,抑制土壤盐分表聚,增加土壤 N、K、P 等养分有效性,提高作物产量。本文分析、归纳了近年来酸性物质对苏打盐碱土理化性质、水盐运移、土壤肥力、作物生长和产量等方面的影响,探讨其作用规律和机制,为苏打盐碱土农业利用和可持续发展提供理论依据。

1 酸性物质改良盐碱土的机理

酸性物质施入土壤能够降低盐碱土 pH,溶解土壤中的 Ca^{2+} ,为代换土壤中的交换性钠提供 Ca^{2+} 源。溶解的 Ca^{2+} 与土壤中的 Na^+ 发生交换吸附反应,从而置换出 Na^+ ,土壤碱化度(ESP)和钠吸附比(SAR)降低,提高土壤入渗率,随着 Ca^{2+} 代换 Na^+ 进入代换位和入渗水的增加, Ca^{2+} 的溶解量和 Na^+ 的淋洗量逐渐增加,土壤的结构和理化性质得到改善,盐碱土逐步得到改良。常见的酸性物质改良盐碱土的机理、优点和缺点如表 1 所示。

总体来看,酸性物质能快速有效地改良苏打盐碱土。苏打盐碱土其自身含有碳酸钙^[22],盐碱土施加酸性物质后,土壤中 H^+ 含量增加,促进土壤中碳酸钙溶解,为土壤中 Na^+ 交换吸附反应提供较多的钙源。但早前人们采用石膏、磷石膏等含钙物质改良盐碱土,通过外源增加土壤中 Ca^{2+} 的活度,将吸附于土壤胶体中的 Na^+ 交换出来。但目前对酸性物质改良苏

表 1 不同酸性物质改良盐碱土的机理及优缺点
Table 1 Mechanisms and advantages of different acidic substances for improving saline soils

改良剂 Amendment	机理 Mechanism	优点 Advantage	缺点 Disadvantage	参考文献 Reference
硫酸 Sulfuric acid	<p>硫酸可解离盐碱土中固有的碳酸钙, 使Ca²⁺置换土壤胶体吸附的Na⁺</p> $H_2SO_4 + CaCO_3 + H_2O \rightarrow CaSO_4 \cdot H_2O + CO_2$ <p>土壤胶体(Na⁺) + CaSO₄ → 土壤胶体 (Ca²⁺) + Na₂SO₄</p> <p>Sulfuric acid dissociates CaCO₃ inherent in saline soils, allowing Ca²⁺ to replace Na⁺ adsorbed on soil colloids!</p> $H_2SO_4 + CaCO_3 + H_2O \rightarrow CaSO_4 \cdot H_2O + CO_2$ $\text{Soil colloid (Na}^+) + CaSO_4 \rightarrow \text{soil colloid (Ca}^{2+}) + Na_2SO_4$	<p>在盐碱土中施用硫酸可以降低土壤pH, 溶解土壤中的碳酸钙, 增加Ca²⁺含量, 使Na⁺从土壤胶体置换到土壤溶液中, 有利于Na⁺随水分排出, 且工业废硫酸价格低廉, 易于购买</p> <p>Application of sulfuric acid in saline soils can lower soil pH, dissolve CaCO₃ in the soil, increase Ca²⁺ content, displace Na⁺ from the soil colloid into the soil solution, facilitate Na⁺ discharge with water, and industrial waste sulfuric acid is cheap and easy to purchase</p>	<p>硫酸腐蚀性很强, 使用不便, 为便于操作, 可先采用固体载体吸附或对硫酸进行改性</p> <p>Sulfuric acid is highly corrosive and inconvenient to use. It can be used after adsorption on solid carrier or after modification</p>	[23-24]
硫酸亚铁 Ferrous sulfate	<p>硫酸亚铁溶于水, Fe²⁺发生水解产生H⁺, 降低土壤pH, 促进碳酸盐溶解, 与土壤胶体上的Na⁺发生交换</p> $Fe^{2+} + 2H_2O \rightarrow 2H^+ + Fe(OH)_2$ <p>Fe²⁺水解生成的氢氧化亚铁, 在碱性环境中极易被氧化为氢氧化铁, 它具有较强的絮凝作用, 使土壤微粒逐渐团聚, 从而增加土壤饱和导水率, 并降低土壤容重</p> $4Fe(OH)_2 + O_2 + 2H_2O \rightarrow 4Fe(OH)_3$ <p>When ferrous sulphate is dissolved in water, Fe²⁺ are hydrolyzed to H⁺, which lowers the pH of the soil, promotes the dissolution of CaCO₃ and its substitutes with Na⁺ on the soil colloid</p> $Fe^{2+} + 2H_2O \rightarrow 2H^+ + Fe(OH)_2$ <p>Ferrous hydroxide, produced by the hydrolysis of Fe²⁺, is easily oxidized to ferric hydroxide in an alkaline environment, which has a strong flocculating effect, causing soil particles to gradually agglomerate, thereby increasing the saturated hydraulic conductivity of the soil and reducing the soil bulk density</p>	<p>硫酸亚铁能够降低盐碱土pH, 降低表层土壤交换性钠含量和土壤的碱化度(ESP), 降低土壤容重, 提高饱和导水率</p> <p>Ferrous sulphate can lower the pH of saline soils, reduce the exchangeable sodium content and exchangeable sodium saturation percentage (ESP) of topsoil, lower the soil bulk density and increase saturated hydraulic conductivity</p>	<p>硫酸亚铁价格较低, 但主要产地在山东等其他地区, 其运输成本高于产品成本, 且单位面积使用量较大, 总成本较高</p> <p>The price of ferrous sulfate is low, but its main production area is in Shandong and other areas, its transportation costs higher than the cost of the product, and the use of a large amount of unit area, the total cost is higher</p>	[25-27]
硫磺 Sulfur	$4Fe(OH)_2 + O_2 + 2H_2O \rightarrow 4Fe(OH)_3$ <p>硫磺在土壤中的氧化过程为:</p> $S \rightarrow S_2O_3^{2-} \rightarrow S_2O_6^{2-} \rightarrow SO_4^{2-} \rightarrow SO_4^{2-}$ <p>产生H⁺从而使土壤pH降低; 与苏打盐碱土中难溶的碳酸钙发生反应, 生成高溶解度的硫酸钙, 提高土壤活性钙含量, 提升土壤改良效果</p> <p>The oxidation process of sulfur in soil is:</p> $S \rightarrow S_2O_3^{2-} \rightarrow S_2O_6^{2-} \rightarrow SO_4^{2-} \rightarrow SO_4^{2-}$ <p>which generates H⁺ and thus lowering soil pH. Sulfur reacts with insoluble CaCO₃ in saline-sodic soils to produce highly soluble CaSO₄, which increases the active calcium content of the soil and enhances the soil improvement effect</p>	<p>硫磺对人畜无毒, 不易对作物产生毒害现象, 此外硫磺可以从燃烧产生的含硫天然气和石油废气中回收利用, 降低含硫废气对环境造成污染</p> <p>Sulfur is harmless to humans and animals, and is not easily toxic to crops. In addition, sulfate can be recycled from sulfater-containing natural gas and petroleum waste gas generated by combustion, reducing the pollution caused by sulfur-containing waste gas to the environment</p>	<p>盐碱土中施用硫磺虽可降低土壤pH, 促进Ca²⁺溶解, 但在施用硫磺时, 要注意用量及其在土壤中的变化趋势, 避免过量施用造成土壤盐害加重, 从而对作物生长造成不利影响</p> <p>Although the application of sulfur in saline soils can reduce soil pH and promote the dissolution of Ca²⁺, it is important to pay attention to the amount of sulfur and its dynamic trend in soil to avoid excessive application of sulfur to aggravate soil salt damage, which may adversely affect crop growth</p>	[28-30]

续表 1

改良剂 Amendment	机理 Mechanism	优点 Advantage	缺点 Disadvantage	参考文献 Reference
聚合硫酸铝铁 Poly-ferric aluminum sulfate	<p>主要成分为$[\text{Al}(\text{OH})_3\text{SO}_4]_m[\text{Fe}_2(\text{OH})_2\text{SO}_4]_n$</p> <p>聚合硫酸铝铁水解可产生大量$\text{H}^+$, 降低土壤pH, 其阳离子的水解产物可以中和带负电荷的土壤胶体</p> <p>The main components are $[\text{Al}(\text{OH})_3\text{SO}_4]_m[\text{Fe}_2(\text{OH})_2\text{SO}_4]_n$</p> <p>It's hydrolysis can produce large amounts of H^+ and lowers soil pH, and the hydrolysis products of its cations neutralize negatively charged soil colloids</p> <p>硫酸铝溶液中Al^{3+}水解产生H^+, 使土壤pH降低, 所产生的H^+与土壤中碳酸盐发生反应, 增加土壤溶液中Ca^{2+}浓度, 从而与土壤胶体中Na^+发生代换, 降低土壤胶体上Na^+含量</p> <p>相关式子如下:</p> <p>铝离子改良剂+H_2O=羟基铝离子+$n\text{H}^+$</p> <p>$2\text{H}^+ + \text{CaCO}_3 = \text{Ca}^{2+} + \text{H}_2\text{O} + \text{CO}_2$</p> <p>土壤胶体-$2\text{Na}^+ + \text{Ca}^{2+}$=土壤胶体-$\text{Ca}^{2+} + 2\text{Na}^+$</p> <p>The hydrolysis of Al^{3+} in the aluminum sulfate solution produces H^+, which lowers soil pH. The produced H^+ reacts with CaCO_3 in soil to increase the Ca^{2+} concentration in soil solution, thus substituting with Na^+ in the soil colloid and reducing Na^+ content on the soil colloid</p> <p>The relevant equations are as follows:</p> <p>Aluminum ion modifier + H_2O = hydroxy aluminum ion + $n\text{H}^+$</p> <p>$2\text{H}^+ + \text{CaCO}_3 = \text{Ca}^{2+} + \text{H}_2\text{O} + \text{CO}_2$</p> <p>Soil colloid-$2\text{Na}^+ + \text{Ca}^{2+}$ = soil colloid-$\text{Ca}^{2+} + 2\text{Na}^+$</p> <p>磷酸脲是由磷酸和尿素等摩尔反应生成, 反应方程式为</p> <p>$\text{H}_3\text{PO}_4 + \text{CO}(\text{NH}_2)_2 \rightarrow \text{CO}(\text{NH}_2)_2\text{-H}_2\text{PO}_4$</p> <p>磷酸脲增加土壤中$\text{H}^+$浓度, 降低土壤pH, 抑制$\text{NH}_4^+ \rightarrow \text{NH}_3 \uparrow + \text{H}^+$反应, 降低土壤对磷的固定作用</p> <p>Urea phosphate is produced by equimolar reaction of phosphoric acid and urea, the reaction equation is:</p> <p>$\text{H}_3\text{PO}_4 + \text{CO}(\text{NH}_2)_2 \rightarrow \text{CO}(\text{NH}_2)_2\text{-H}_2\text{PO}_4$</p> <p>Phosphate urea increases H^+ concentration in soil, lowers soil pH, inhibits $\text{NH}_4^+ \rightarrow \text{NH}_3 \uparrow + \text{H}^+$ reaction, and reduces phosphorus fixation by soil</p>	<p>可以促进土壤团聚, 改善土壤孔隙度、持水能力和渗透性, 促进土壤碳酸盐溶解</p> <p>Poly-ferric aluminum sulfate can promote soil agglomeration, improve soil porosity, water holding capacity and permeability, and promote dissolution of CaCO_3 in soil</p> <p>铝离子水解产物可以促进土壤胶体凝聚, 进而增加土壤微生物数量。同时, 硫酸铝还可降低土壤容重, 增加土壤孔隙度, 改善和提高土壤性质</p> <p>Aluminum ion hydrolysis products can promote soil colloid coalescence, thus increasing the number of soil microaggregates. Aluminum sulfate can also reduce soil bulk density, increase soil porosity, improve and enhance soil properties</p>	<p>与硫酸铝相比, 聚合硫酸铝铁对水溶性钾离子含量的影响很小, 并且对苏打盐碱土中阳离子交换量影响较小</p> <p>Compared to aluminum sulfate, Poly-ferric aluminum sulfate has little effect on the water-soluble potassium ion content and less effect on cation exchange in soda saline soils</p>	[31]
硫酸铝 Aluminum sulfate	<p>铝离子改良剂+H_2O=羟基铝离子+$n\text{H}^+$</p> <p>$2\text{H}^+ + \text{CaCO}_3 = \text{Ca}^{2+} + \text{H}_2\text{O} + \text{CO}_2$</p> <p>土壤胶体-$2\text{Na}^+ + \text{Ca}^{2+}$=土壤胶体-$\text{Ca}^{2+} + 2\text{Na}^+$</p> <p>The hydrolysis of Al^{3+} in the aluminum sulfate solution produces H^+, which lowers soil pH. The produced H^+ reacts with CaCO_3 in soil to increase the Ca^{2+} concentration in soil solution, thus substituting with Na^+ in the soil colloid and reducing Na^+ content on the soil colloid</p> <p>The relevant equations are as follows:</p> <p>Aluminum ion modifier + H_2O = hydroxy aluminum ion + $n\text{H}^+$</p> <p>$2\text{H}^+ + \text{CaCO}_3 = \text{Ca}^{2+} + \text{H}_2\text{O} + \text{CO}_2$</p> <p>Soil colloid-$2\text{Na}^+ + \text{Ca}^{2+}$ = soil colloid-$\text{Ca}^{2+} + 2\text{Na}^+$</p> <p>磷酸脲是由磷酸和尿素等摩尔反应生成, 反应方程式为</p> <p>$\text{H}_3\text{PO}_4 + \text{CO}(\text{NH}_2)_2 \rightarrow \text{CO}(\text{NH}_2)_2\text{-H}_2\text{PO}_4$</p> <p>磷酸脲增加土壤中$\text{H}^+$浓度, 降低土壤pH, 抑制$\text{NH}_4^+ \rightarrow \text{NH}_3 \uparrow + \text{H}^+$反应, 降低土壤对磷的固定作用</p> <p>Urea phosphate is produced by equimolar reaction of phosphoric acid and urea, the reaction equation is:</p> <p>$\text{H}_3\text{PO}_4 + \text{CO}(\text{NH}_2)_2 \rightarrow \text{CO}(\text{NH}_2)_2\text{-H}_2\text{PO}_4$</p> <p>Phosphate urea increases H^+ concentration in soil, lowers soil pH, inhibits $\text{NH}_4^+ \rightarrow \text{NH}_3 \uparrow + \text{H}^+$ reaction, and reduces phosphorus fixation by soil</p>	<p>可以促进土壤团聚, 改善土壤孔隙度、持水能力和渗透性, 促进土壤碳酸盐溶解</p> <p>Poly-ferric aluminum sulfate can promote soil agglomeration, improve soil porosity, water holding capacity and permeability, and promote dissolution of CaCO_3 in soil</p> <p>铝离子水解产物可以促进土壤胶体凝聚, 进而增加土壤微生物数量。同时, 硫酸铝还可降低土壤容重, 增加土壤孔隙度, 改善和提高土壤性质</p> <p>Aluminum ion hydrolysis products can promote soil colloid coalescence, thus increasing the number of soil microaggregates. Aluminum sulfate can also reduce soil bulk density, increase soil porosity, improve and enhance soil properties</p>	<p>与硫酸铝相比, 聚合硫酸铝铁对水溶性钾离子含量的影响很小, 并且对苏打盐碱土中阳离子交换量影响较小</p> <p>Compared to aluminum sulfate, Poly-ferric aluminum sulfate has little effect on the water-soluble potassium ion content and less effect on cation exchange in soda saline soils</p>	[32-37]
磷酸脲 Urea phosphate	<p>磷酸脲是由磷酸和尿素等摩尔反应生成, 反应方程式为</p> <p>$\text{H}_3\text{PO}_4 + \text{CO}(\text{NH}_2)_2 \rightarrow \text{CO}(\text{NH}_2)_2\text{-H}_2\text{PO}_4$</p> <p>磷酸脲增加土壤中$\text{H}^+$浓度, 降低土壤pH, 抑制$\text{NH}_4^+ \rightarrow \text{NH}_3 \uparrow + \text{H}^+$反应, 降低土壤对磷的固定作用</p> <p>Urea phosphate is produced by equimolar reaction of phosphoric acid and urea, the reaction equation is:</p> <p>$\text{H}_3\text{PO}_4 + \text{CO}(\text{NH}_2)_2 \rightarrow \text{CO}(\text{NH}_2)_2\text{-H}_2\text{PO}_4$</p> <p>Phosphate urea increases H^+ concentration in soil, lowers soil pH, inhibits $\text{NH}_4^+ \rightarrow \text{NH}_3 \uparrow + \text{H}^+$ reaction, and reduces phosphorus fixation by soil</p>	<p>磷酸脲中磷和氮的总养分含量可达60%以上, 可减少土壤挥发性氮损失, 增加土壤养分</p> <p>The total content of phosphorus and nitrogen in urea phosphate is more than 60%, which reduces volatile losses of soil nitrogen, and increases soil nutrients contents</p>	<p>磷酸脲的营养元素含量不均, 仅含有氮和磷, 在市场上仅作为肥料添加剂使用, 导致肥料利用率下降, 限制了其广泛应用</p> <p>Urea phosphate has uneven nutrient content, containing only nitrogen and phosphorus, and is only used commercially as a fertilizer additive, resulting in lower fertilizer utilization and limiting its wide application</p>	[38-40]

打盐碱土的机理了解较少,如硫酸、磷酸脲等酸性物质很少应用于改良盐碱土,其广泛应用存在问题。

2 酸性物质对盐碱土物理性质的影响

土壤的物理性质主要包括土壤的容重、孔隙度、团聚体和土壤持水特性等。土壤盐度和碱化度共同作用导致钠饱和黏土分散^[41],破坏土壤结构,使土壤颗粒分散,潮湿时泥泞透气性差,孔隙度低,干时紧实板结耕性极差,对作物的生长发育造成不利影响^[42-43]。

2.1 土壤团聚体

苏打盐碱土中的盐分主要为苏打(碳酸钠, NaCO_3)、小苏打(碳酸氢钠, NaHCO_3), Na^+ 含量高,导致土壤结构性差、养分匮乏以及微生物活性较低。大量 Na^+ 会导致土壤结构崩塌,破坏土壤团聚体的稳定性^[44]。土壤团聚体是土壤结构的基本单位,其质量与数量对土壤性质起决定性作用^[45];

研究发现,在重度苏打盐碱土中添加酸性物质如硫酸铝改良剂后,土壤溶液 pH 显著下降。在一定范围内,土壤中大部分微团聚体含量随硫酸铝添加量的增加而增加,大粒径团聚体数量明显增多,孔隙度增大,容重质量减小,作物生长状况得到改善,出苗率提高^[46]。单纯施用硫酸铝虽可有效改善土壤结构,但过去的研究表明^[47-48],施用家畜粪、秸秆等有机物质有利于土壤形成大颗粒团聚体,提高团聚体稳定性。建议在生产实践中,可将硫酸铝与其他有机物料配合使用,可进一步改善土壤结构,提高土壤肥力。近几年来,有学者通过用中低浓度的盐酸或硫酸对沸石进行酸化改性进而改良苏打盐碱土^[49]。对沸石进行硫酸酸化改性后,硫酸能有效地去除沸石中金属离子, H^+ 取代金属离子,增强沸石吸附能力,使土壤颗粒在沸石颗粒周围形成较为稳定的胶体粒子,不会再分散。因此,在苏打盐碱土中施用酸改性沸石,可以改善土壤中 Na^+ 含量过高与土壤板结的问题。

2.2 容重与孔隙度

对盐碱土而言,不同土壤孔径的组成与分布直接影响土壤结构,而土壤孔隙则能真实反映土壤结构状况^[50]。土壤的容重和孔隙度是土壤的重要物理性质,它直接影响土壤水分、肥力等因素变化。

研究表明^[51],硫酸铝可以显著降低苏打盐碱土旱田和水田土壤容重,改善土壤微团聚体组成,使土壤结构状况得到明显改善。聚合硫酸铝铁施用到土壤

中,其阳离子的水解产物能与土壤中带负电荷的胶体进行电荷中和,促进土壤团聚结构的形成,改善土壤孔隙度^[52]。目前仅有少量研究表明,硫酸也可用于改良盐碱土。硫酸能降低土壤 pH 和盐分含量,改善土壤环境。硫酸可溶解土壤中的碳酸钙,通过用钙代替土壤中的交换性钠来提高土壤的孔隙度,降低土壤的紧实度^[53]。

2.3 渗透性与持水性

土壤渗透性和持水性决定着土壤淋洗排盐效率^[54]。土壤入渗是田间水循环的重要组成部分,酸性物质如硫酸铝可以有效改善苏打盐碱土的土壤结构性与持水特性,并且随着硫酸铝用量的增加,土壤的膨胀性、渗透性和持水性都会增强^[55]。通过室内垂直淋洗试验发现^[56],聚合硫酸铝铁能提高苏打盐碱土的导水率,促进土壤淋洗脱盐。目前苏打盐碱土改良大多采用石膏等含钙物质,通过添加外源钙置换土壤中的 Na^+ 。但 Ahmad 等^[25]通过硫酸和石膏对盐碱土壤离子置换的试验结果表明,硫酸和石膏均可降低土壤饱和泥浆电导率(EC)和 SAR,且经过硫酸处理的土壤中盐分和 Na^+ 的浸出量与石膏的浸出量相当,可能是由于硫酸与碳酸钙反应释放 Ca^{2+} ,在置换土壤胶体上吸附的 Na^+ 所需时间较长^[57]。由于这种有效但延迟浸出的盐和离子,经硫酸处理的土壤在试验结束时相对保持了更多的 EC, SAR 值更低,可以更好地疏松土壤,增加土壤的渗透性。

酸性物质对苏打盐碱土物理性质的改良作用主要有以下几个方面: 1) 凝聚土壤颗粒,改善土壤结构。某些酸性物质具有膨胀性、分散性、黏着性等特性,能使土壤中因盐碱而分散的土壤颗粒凝结,从而改变土壤的孔隙度,改善土壤的通透性。2) 酸性物质可改善土壤微团聚体的组成,增加土壤中大粒径团聚体的数量,增加土壤孔隙度,降低土壤容重,提高土壤淋洗速率。3) 硫酸等强酸性物质可以与盐碱土中的碳酸钠发生反应,使土壤 Ca^{2+} 活化,进而交换土壤胶体吸附的 Na^+ ,阻止胶体分散,增加土壤颗粒絮凝和胶结,改善土壤结构。

3 酸性物质对土壤盐碱性质的影响

衡量盐碱土壤盐碱化程度的主要指标有 pH、电导率(EC)、ESP 等。盐碱土壤中含有大量碱性盐类物质,在碱性环境中容易发生盐碱化。盐碱土中的碳酸钠和碳酸氢钠易发生水解,离解生成 HCO_3^- 、 OH^- ,土壤中含有大量可溶盐类(如 Na^+ 和 Mg^{2+}),这些盐离子进入土壤后容易与土壤中的物质结合形成复

合物,与土壤中游离的 HCO_3^- 、 OH^- 发生反应,从而使土壤呈碱性^[58]。

目前已有不少学者致力于研究硫酸铝对苏打盐碱土的改良效果。硫酸铝改良剂能显著降低苏打盐碱土 pH,并能显著提高土壤中 Ca^{2+} 、 Mg^{2+} 、 K^+ 、 Na^+ 等浓度。在室内模拟试验中^[59],随着硫酸铝用量的增加,土壤阳离子交换量(CEC)呈递增趋势,交换性 Na^+ 含量和 ESP 随硫酸铝用量的增加而逐渐降低。硫酸铝施加到新开垦苏打盐碱水田中,与未处理土壤相比,土壤 pH 值、全盐量、交换性钠、阳离子交换量及碱化度均有所下降;其中 CO_3^{2-} 、 Ca^{2+} 、 K^+ 、 Na^+ 的含量较低, SO_4^{2-} 含量较高^[60]。

盐碱土施用硫酸亚铁后,硫酸亚铁能够水解出 H^+ ,降低土壤 pH^[61],促进土壤碳酸钙的溶解,使土壤中的 Ca^{2+} 与土壤胶体上的 Na^+ 发生代换,降低土壤 ESP^[62]。范定慷^[63] 通过不同含硫量物质对苏打盐碱土的改良效应及盐分运移规律进行了试验研究,结果表明:硫酸亚铁、硫酸铝、硫磺均能降低土壤 pH、SAR、全盐量、ESP 等。

硫磺经微生物作用或水解后,会产生酸性物质,中和盐碱土中的碱性物质(如碳酸钠和碳酸氢钠),进一步降低土壤 pH,改善土壤结构,促进作物生长。如果施用于石灰性土壤,会与土壤中难溶解的碳酸钙作用后生成溶解度大的硫酸钙,土壤中活性钙含量增加,改良作用也随之增强^[30]。刘刚等^[64] 在盐碱土改良土壤中施用硫磺,发现硫磺处理后土壤 pH 下降,部分交换性钠离子含量降低, Ca^{2+} 、 Mg^{2+} 、 SO_4^{2-} 含量增加。已有研究表明^[65-68],硫酸可以迅速降低土壤 pH,特别是在 ESP 较高的土壤中,硫酸与盐碱土中的碳酸钙发生反应,增加可溶性钙离子含量,加快钙钠离子交换速率。但硫酸对苏打盐碱土的应用较少,其对苏打盐碱土交换性钠降低速率尚不明确。

酸性物质对苏打盐碱土盐碱性质的作用机制主要有:1)降低土壤 pH。酸性物质施入土壤后,可降低盐碱土 pH 值,进而溶解和提高土壤中钙离子含量,并与土壤中钠离子发生反应,土壤的 ESP 和 SAR 随之降低。2)影响土壤可溶性盐组成。如盐碱土中加入硫酸铝后,土壤的盐分组成发生变化,由苏打型→苏打-硫酸盐型→硫酸盐型。

4 酸性物质对盐碱土养分的影响

土壤养分是作物生长发育所必需的营养元素,盐碱土因其理化特性而引发离子拮抗和渗透胁迫^[69],从而导致土壤通常缺乏氮磷钾等营养元素,且有机质溶解度较高,养分淋失加快。植物根系活性降低,

对土壤养分转换和效率产生一定的不利作用。抑制植物吸收利用矿质养分和水分。土壤养分均衡可以为作物提供良好的生长环境;但养分失衡,则会对作物的生长发育产生不利作用,导致作物的产量下降和病害加重^[70]。

4.1 有机质

土壤有机碳在全球碳循环中起着重要作用,也是土壤肥力的重要组成部分。土壤有机碳是评价土壤肥力和质量的重要指标,对改善土壤物理特性、提供养分、保护环境具有重要作用^[71-72]。活性有机碳是土壤有机质有效组分之一,它能反映土壤有机质的微小变化^[73],直接影响植物的养分供应^[74-75]。

采用不同比例硫酸铝、稻草和苏打盐碱土进行室内模拟培养试验,结果表明^[76]:在一定范围内,单独施用硫酸铝可增加土壤易氧化有机碳含量,降低土壤有机碳的氧化稳定系数。硫酸铝和稻草同时施用促进秸秆分解,进一步增加土壤有机质。张鑫^[77] 同样进行室内模拟试验,结果表明:苏打盐碱土微团聚体中有机碳含量在添加硫酸铝改良剂和玉米秸秆后明显增加,土壤大粒级微团聚体的有机碳含量和贡献率提高,有利于提高有机碳固定性能。

硫酸铝单独施用可增加土壤有机质含量,而秸秆还田和添加有机肥则能进一步提高土壤有机质含量^[78-79]。然而,目前较多学者主要探究硫酸铝对苏打盐碱土有机质的影响,硫酸、硫磺及硫酸铝铁等酸性物质的研究较少,可先通过室内模拟试验,探究其他酸性物质对苏打盐碱土有机质的影响,再应用到大田试验中。

4.2 营养元素

苏打盐碱土碱性较高,土壤自然肥力较差,在碱性环境下,土壤中磷钾等营养元素被固定^[80-81],限制了土壤养分有效性^[82-85]。营养元素缺乏和盐碱胁迫是制约作物生产的两个主要问题。在盐碱条件下,土壤高 pH、高电导率、有机质含量低等因素,导致土壤养分含量下降,对作物生长不利^[86-87]。在苏打盐碱土中施加硫酸、硫酸铝等酸性物质可有效降低土壤 pH,提高土壤中磷钾等元素有效性^[88]。

已有研究^[89]表明,硫酸铝配施磷肥和氮肥均能显著促进小麦(*Triticum aestivum*)出苗和生长,无论是否施加磷肥,土壤有效磷曲线均随着硫酸铝用量的增加而逐渐降低。马巍等^[19]对苏打盐碱地水稻(*Oryza sativa*)施用硫酸铝,结果表明:在盐碱化严重的稻田中,施用硫酸铝可以提高水稻吸收氮磷钾的能力,提高吸磷量,增加磷肥利用率。磷酸脲在土壤

中分解成尿素和磷酸, 磷素在土壤中所占比例较高, 从而增加了土壤磷素供给强度。磷酸脲在苏打盐碱土中的应用较少, 但已有研究^[90]表明, 磷酸脲对土壤中磷铁锰等元素的活化具有显著的促进作用, 为作物生长发育提供更多的营养元素, 提高作物产量。杨贵婷^[38]以磷酸脲基复合肥为改良剂, 进行为期两年的玉米 (*Zea mays*) 田间试验, 结果表明: 经全量磷酸脲基复合肥处理的玉米产量、磷素利用率均高于普通磷肥。磷酸脲作为改良剂施入苏打盐碱土中, 可能会降低土壤 pH, 改善土壤理化性质, 增加土壤磷、钙等元素含量。

5 酸性物质对盐碱土作物生长及产量的影响

苏打盐碱土含有碳酸钠、重碳酸钠等盐分, 影响土壤物理、化学性质以及微生物活性, 不仅会抑制作物生长, 还会导致作物死亡, 生产力下降^[91-92]。研究表明^[93], 酸性物质可改善土壤的理化性质, 促进作物生长。

硫酸亚铁是一种工业副产品, 其利用率低, 导致大量硫铁资源浪费。硫酸亚铁可以作为铁肥、硫肥、调酸剂等对苏打盐碱土进行改良, 可提高水稻的抗病能力, 提高秧苗素质。经硫酸亚铁培养的水稻秧苗, 其体内的铁硫含量增加; 叶绿素含量增多, 秧苗素质提高^[94-95]; 硫酸亚铁对苏打盐碱土水稻分蘖有明显的促进作用, 有效穗数、单穗粒数和产量均有明显提高^[96]。硫酸铝水解产生的 H^+ 可与土壤中 CO_3^{2-} 和 HCO_3^- 发生反应生成 H_2O 和 CO_2 , 使碳酸钠和碳酸氢钠变成硫酸钠, 土壤类型转变为硫酸盐型, 减轻碳酸钠和碳酸氢钠对作物生长的危害, 对苏打盐碱土水稻生长环境有一定的优化作用, 为水稻秧苗的生长提供了有利的环境条件, 提高了水稻产量^[97-98]。硫酸除有效降低土壤 pH 外, 还能提高水稻的生理适应能力和产量, 苏打盐碱土施用浓硫酸可显著改善土壤理化性状和水稻的生长环境, 增加水稻分蘖, 并提高水稻的单株有效穗数、单株实粒数、千粒重、结实率, 进而增加水稻产量^[99]。

酸性物质提高苏打盐碱土作物生长及产量的可能原因有: 1) 酸性物质可降低土壤 pH, 活化土壤中的碳酸钙, 增加土壤 Ca^{2+} 含量, 解析盐碱土中的 Na^+ , 从而降低高浓度 Na^+ 对作物生长的危害, 促进作物生长。2) 酸性能够降低苏打盐碱土 pH, 提高养分有效性, 增加土壤肥力。3) 酸性物质可降低盐碱土中有毒元素富集, 降低土壤中原有重金属对作物生长的影响。

6 展望及未来发展方向

苏打盐碱土的改良是一个持续性过程, 目前众多改良方法都存在着投资大、见效慢等问题。利用酸性物质改良苏打盐碱土, 相对其他方法而言见效快且方便。酸性物质可以改善土壤理化性质、提高土壤的透水透气能力和入渗速率以及增加土壤中微量元素的有效性, 进而提升土壤肥力, 增加农作物产量, 是提高农业生产效率和环境可持续发展的重要因素。由于不同地区农业气候资源、土壤水肥气热条件及农业管理模式差异较大, 因此在实际应用酸性物质时应因地制宜, 综合不同气候、土壤、作物等因素, 确定适宜的改良剂用量和施用方法, 提高苏打盐碱土改良效果。

目前, 酸性物质在苏打盐碱土的改良应用中还存在一些亟待解决的问题。

1) 活化土壤固有“钙源”, 减少外源钙质材料添加。苏打盐碱土本身含有钙源物质, 即碳酸钙, 但其溶解度较低, Ca^{2+} 较难析出。而添加酸性物质可以使土壤中 H^+ 含量增加, 促进碳酸钙溶解, 增加土壤 Ca^{2+} 含量, 使 Ca^{2+} 置换土壤胶体上吸附的 Na^+ 。这种改良方法可以减少外源钙质改良剂的添加, 利用其土体本身的“ Ca^{2+} ”对土壤进行改良。在苏打盐碱土未来改良中, 可以选择酸性物质对其进行改良, 并探究不同酸性物质对苏打盐碱土的改良机理和最佳施用量。在选择酸性改良剂时, 可选用一些工业废液, 如工业废硫酸、硫酸铝、硫酸亚铁等, 在降低改良成本的同时, 也可减少工业废弃物对环境造成污染。

2) 改良剂“多元化”发展。较多学者大多采用硫酸铝和硫酸亚铁对苏打盐碱土进行改良, 硫酸、磷酸脲等酸性物质应用较少。在苏打盐碱土未来改良中, 可选用其他酸性物质, 探究其对苏打盐碱土有机碳、土壤生物肥力及对作物生长的影响。

3) 选择“绿色”改良剂。酸性物质虽可改良盐碱土, 改善土壤的结构和理化性质, 但需要注意的是, 土壤酸化后会引发土壤重金属离子的活化及其在作物中的富集作用。农作物生产必须以安全和健康为前提。选择酸性物质改良苏打盐碱土需具有以下特点: ①其自身不含重金属超标等有毒物质。②所选用的改良剂不能使土壤产生有毒有害物质。③经改良后的土壤排出物不会污染周边环境^[100]。

4) 增加大田试验。酸性物质对苏打盐碱土的改良多用于室内模拟试验, 但不同类型盐碱土的理化性质十分复杂。应适当增加大田试验, 对酸性物质应用于大田的适用性和作用机理进行分析, 并对其

持续效应进行评估。

5) 选择性价比较高的改良剂。目前我国农业产业效益低,特别是大田作物收益更低。在盐碱土改良时,应选择具有较高性价比的改良剂,促进苏打盐碱土改良和农业利用持续发展。

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